

Atomic Symbol Worksheet

Name: _____ Date: _____ Period: _____



- * Atomic number is always the same as number of protons.
- * # of protons = # of electrons unless you have an ion (+ or - charge such as Cu^{+2})
- * Mass number (atomic mass rounded off) = protons + neutrons
- * When a symbol in a problem is written as: ${}^{14}_6\text{C}$, the top number is the mass and the bottom is the atomic number.

Fill in missing information:

Symbol	Atomic number	Number of protons	Mass number	Number of neutrons	Number of electrons
${}^1_1\text{H}$					
H	1			1	
		1	3		
	6			8	
		6	13		
${}^{12}_6\text{C}$					
${}^{90}_{38}\text{Sr}$					
	16		34		
${}^{32}_{16}\text{S}$					
		55	137		
${}^{239}_{94}\text{Pu}$					
${}^{235}_{92}\text{U}$					
	92			146	

Inside the atom

Atom (A)/ Ion(I)	Symbol	Name	Atomic Number	Atomic Mass	# of Protons	# of Neutrons	# of Electrons	Overall Charge
	N							
		Neon						
	Ba							
					79			
		Lead						
			20					
					13		10	
					13		13	
			94					
			1			0	1	
			1			1	1	
			1			2	1	
		Uranium						
	Mg^{+2}							
	Li^{+1}							
	Br^{-1}							